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20 4 2-) Example complex N O [Cr(NH 2CH 2CH 2NH 2) 3] 3+ [Cr(C 2O 4) 3] 3- Multidentate ligands – ligands which form more than two co-ordinate bonds to a metal ion.Ligand EDTA4- forms 6 bonds porphyrin forms 4 bonds How it bonds Example [Cu(EDTA)]2- haemoglobin CH2 CH2 H2N :NH2-O: C:O-O O NH 2 Cr NH 2 H 2N NH 2 H 2 NH 2 3+ O Cr NH2 O O O O C ...

### Chemsheets A2 1027 (Transition Metals)

Decomposition chemical reaction is the reaction where only one compound decomposes and results in two or more than two products. Pb(NO 3) 2 PbO + NO 2 + O 2. In this equation, lead nitrate is being decomposed, which breaks down to form nitrogen dioxide, oxygen, and lead oxide. This is an example of a decomposition reaction.

### 49 Balancing Chemical Equations Worksheets [with Answers]

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Example 3 50.0 cm<sup>3</sup> of 0.500 mol dm<sup>-3</sup> copper sulphate solution was put in a calorimeter and 2.0 g of zinc powder added. The temperature of the solution rose by 24.0 C.

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Full worked solutions in: Chemsheets AS029 (Calorimetry) TASK 2 - Calorimetry calculations 2 1 -3800 kJ mol-1 2 -51.8 kJ mol-1 3 -2500 kJ mol-1 4 +20.9 kJ mol-1 5 -4203 kJ mol-1 6 -3920 kJ mol-1 7 -54.3 kJ mol-1 8 -1145 kJ mol-1 9 -1223 kJ mol-1 10 -154 kJ mol-1 11 exo, exo, endo 12 -503 kJ mol-1

### Chemsheets AS 029 (Thermodynamics) ANS 2

n is called the order of reaction with respect to that reactant, and is usually 2, 1 or 0 Order Zero order First order Second order n 0 1 2 The effect the reaction rate is proportional the reaction rate is proportional to [A]<sup>0</sup> (= 1) and so the rate is not affected by changes in [ A] to [A]<sup>1</sup> to [A]<sup>2</sup> [A] x 2 rate unchanged rate x 22 (i.e. x 4)

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2,2-dimethylamine 2-methylbut-1-ene 3-methylbutan-1-ol 4-hydroxybutanone 3-methylbutanoic acid butyl methanoate ethanenitrile propanal Decide which spectrum belongs to which compound and draw the molecule next to the spectrum.

### Chemsheets A2 029 (Spectroscopy) - Chemistry Home page

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### SECTION 1 1H NMR - Weebly

Example 1 Imagine a reaction where R reacts with S to make some products: R + S → products If we compare experiments 1 and 2: Initial [R] Experiment mol dm<sup>-3</sup> 1.0 2.0 Init-ai [S] mo dm<sup>-3</sup> 1.0 1.0 2.0 The

### Chemsheets Kinetics 1 Answers

the value of -114.2 kJ mol<sup>-1</sup> in the first equation means that 114.2 kJ of heat energy is released when 1 mole of H 2 SO 4 reacts with 2 moles of NaOH, the value of -57.1 kJ mol<sup>-1</sup> in the second equation means that 57.1 kJ of heat energy is released when ½ mole of H 2 SO 4 reacts with 1 mole of NaOH.

### Chemsheets a2 1014 thermodynamics booklet - BIOC1008 - UCL ...

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